



OWNER'S MANUAL

OXYGEN SENSOR

Models: SO-110, SO-120, SO-210, SO-220



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DECLARATION OF CONFORMITY

CE and ROHS Certificate of Compliance

We Apogee Instruments, Inc.
721 W 1800 N
Logan, Utah 84321
USA

Declare under our sole responsibility that the products:

Models: SO-110, SO-120, SO-210, SO-220
Type: Oxygen Sensor

are in conformity with the following standards and relevant EC directives:

Emissions: EN 61326-1:2013
Immunity: EN 61326-1:2013
Safety: EN 61010-1:2010

EU directive 2004/108/EC, EMC
EU directive 2002/95/EC, RoHS (Restriction of Hazardous Substances)
EU directive 2011/65/EU, RoHS2

Please be advised that based on the information available to us from our raw material suppliers, the products manufactured by us do not contain, as intentional additives, any of the restricted materials, including cadmium, hexavalent chromium, lead, mercury, polybrominated biphenyls (PBB), polybrominated diphenyls (PBDE).

Further note that Apogee Instruments does not specifically run any analysis on our raw materials or end products for the presence of these substances, but rely on the information provided to us by our material suppliers.

Bruce Bugbee
President
Apogee Instruments, Inc.
June 2013

INTRODUCTION

Oxygen (O₂) is the second most abundant gas in the atmosphere and is essential to life on Earth. Absolute oxygen concentration determines the rate of many biological and chemical processes. Oxygen is required for aerobic respiration. In addition to measurements of absolute oxygen concentration, relative oxygen concentration is also often measured and reported.

Oxygen sensors are used to measure gaseous or dissolved oxygen. There are multiple different techniques for measuring gaseous oxygen. Three of the more widely used sensors for environmental applications are galvanic cell sensors, polarographic sensors, and optical sensors. Galvanic cell and polarographic sensors operate similarly, by electrochemical reaction of oxygen with an electrolyte to produce an electrical current. The electrochemical reaction consumes a small amount of oxygen. Unlike polarographic oxygen sensors, galvanic cell sensors are self-powered and do not require input power for operation. Optical oxygen sensors use fiber optics and a fluorescence method to measure oxygen via spectrometry.

Typical applications of Apogee oxygen sensors include measurement of oxygen in laboratory experiments, monitoring gaseous oxygen in indoor environments for climate control, monitoring of oxygen levels in compost piles and mine tailings, and determination of respiration rates through measurement of oxygen consumption in sealed chambers or measurement of oxygen gradients in soil/porous media.

Apogee Instruments SO-100 and SO-200 series oxygen sensors consist of a galvanic cell sensing element (electrochemical cell), Teflon membrane, reference temperature sensor (thermistor or thermocouple), heater (located behind the Teflon membrane), and signal processing circuitry mounted in a stainless steel housing (or ABS plastic housing for use in acidic environments), and lead wires to connect the sensor to a measurement device. Sensors are designed for continuous gaseous oxygen measurement in ambient air, soil/porous media, sealed chambers, and in-line tubing (flow through applications). SO-100 and SO-200 series oxygen sensors output an analog voltage that is directly proportional to gaseous oxygen.

SENSOR MODELS

Apogee Instruments offers four models of oxygen sensors:

SO-110: slow response, thermistor reference temperature sensor

SO-120: slow response, type-K thermocouple reference temperature sensor

SO-210: fast response, thermistor reference temperature sensor

SO-220: fast response, type-K thermocouple reference temperature sensor

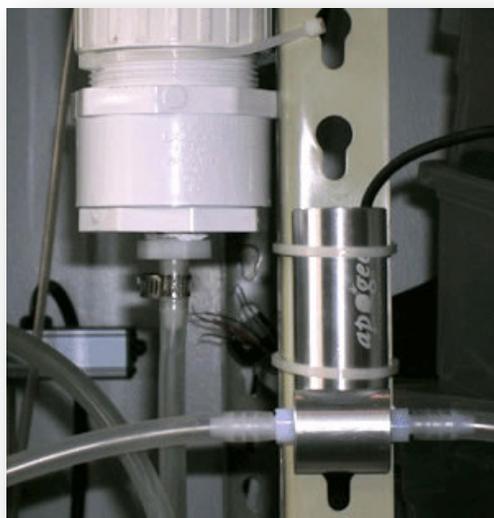
The slow response sensor (SO-100 series) is designed for use in soil/porous media. It has a much longer expected lifetime than the fast response sensor (SO-200 series), which is designed for use in flow through applications.



Sensor model number, serial number, and production date are located on a label between the sensor and pigtail lead wires.



Model SO-110



Model SO-210

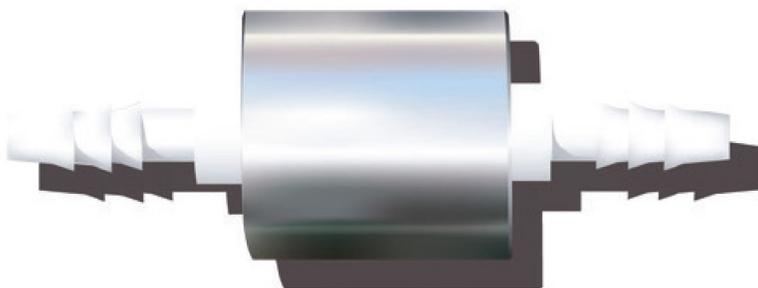
Accessories:

All Apogee oxygen sensors can be purchased with attachments to facilitate measurements in soil/porous media or in-line tubing.

Model AO-001: Diffusion head designed for measurements in soil/porous media. The diffusion head maintains an air pocket and provides protection to the permeable Teflon membrane where gas diffusion occurs.



Model AO-002: Flow through head designed for in-line measurements. The flow through head allows connection of tubing via ¼ inch barbed nylon connectors.



SPECIFICATIONS

Measurement Range: 0 to 100 % O₂

Sensitivity: 2.6 and 0.6 mV per % O₂ for SO-100 and SO-200 series, respectively (approximations)

Output at 0 % O₂: 6 % and 3 % of output at 20.95 % O₂ for SO-100 and SO-200 series, respectively

Measurement Repeatability: < 0.1 % of mV output at 20.95 % O₂

Non-linearity: < 1 %

Non-stability (Signal Decrease): 1 and 0.8 mV per year for SO-100 and SO-200 series, respectively

Oxygen Consumption Rate: 2.2 μmol O₂ per day at 20.95 % O₂ and 23 C (galvanic cell sensors consume O₂ in a chemical reaction with the electrolyte, which produces an electrical current)

Response Time: 60 and 14 s for SO-100 and SO-200 series, respectively (time required to read 90 % of saturated response)

Operating Environment: -20 to 60 C

0 to 100 % relative humidity (non-condensing)

60 to 140 kPa

Note: Electrolyte will freeze at temperatures lower than -20 C. This will not damage the sensor, but the sensor must be at a temperature of -20 C or greater in order to make measurements.

Input Voltage Requirement: 12 V DC continuous (for heater); 2.5 V DC excitation (for thermistor)

Heater Current Drain: 6.2 mA

Thermistor Current Drain: 0.1 mA DC at 70 C (maximum, assuming input excitation of 2.5 V DC)

Dimensions: 3.2 cm diameter, 6.8 cm length

Diffusion Head (Accessory): 3.5 cm diameter, 3.5 cm length, 125 mesh screen

Flow Through Head (Accessory): 3.2 cm diameter, 9.1 cm length, ¼ inch barbed nylon connectors

Mass: 175 g (with 5 m of lead wire)

Cable: 5 m of six conductor, shielded, twisted-pair wire

Additional cable available in multiples of 5 m

Santoprene rubber jacket (high water resistance, high UV stability, flexibility in cold conditions)

Pigtail lead wires

Influence from Various Gases:

Sensors are unaffected by CO, CO₂, NO, NO₂, H₂S, H₂, and CH₄. There is a small effect (approximately 1 %) from NH₃, HCl, and C₆H₆ (benzene). Sensors are sensitive to SO₂ (signal responds to SO₂ in a similar fashion to O₂). Sensors can be damaged by O₃.

DEPLOYMENT AND INSTALLATION

Apogee SO-100 and SO-200 series oxygen sensors are built with an anodized aluminum housing (or ABS plastic housing for use in acidic environments) and are designed to be installed in soil/porous media or sealed chambers, in addition to air. To facilitate the most stable readings, sensors should be mounted vertically, with the opening pointed down and the cable pointed up. This orientation allows better contact between the electrolyte and signal processing circuitry.

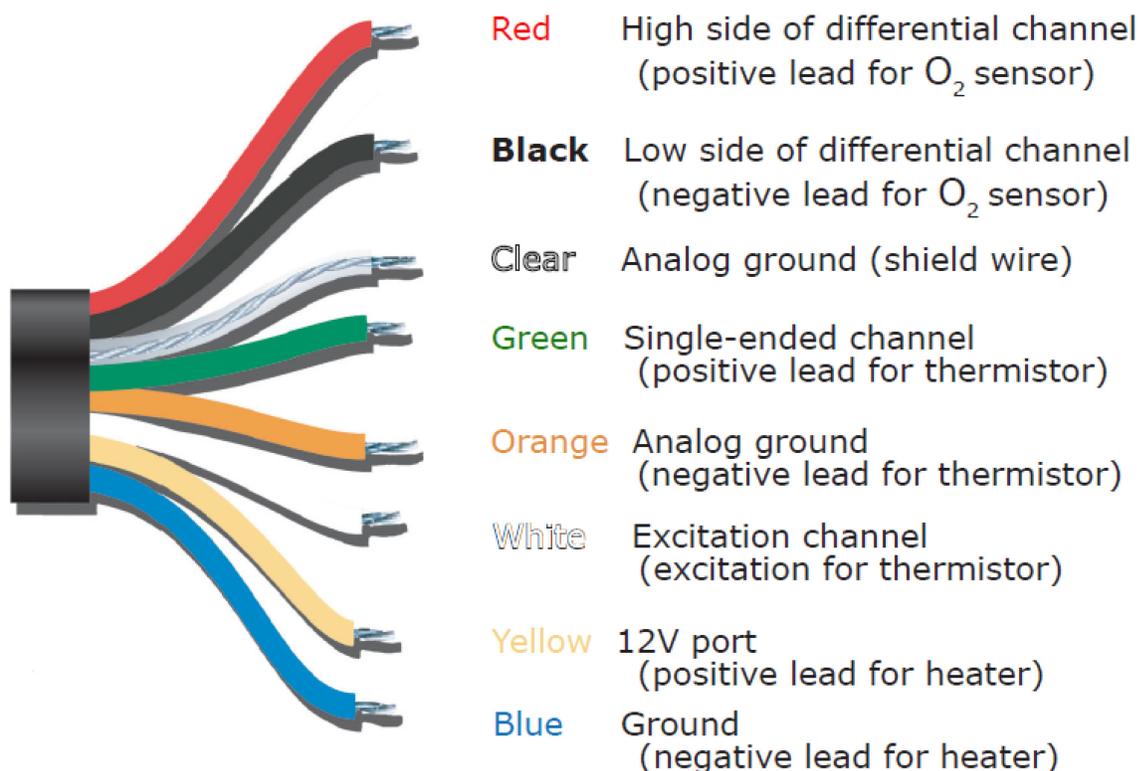


Apogee oxygen sensors are resistant to 2.7 G of shock, but vibration may influence sensor sensitivity and should be minimized.

OPERATION AND MEASUREMENT

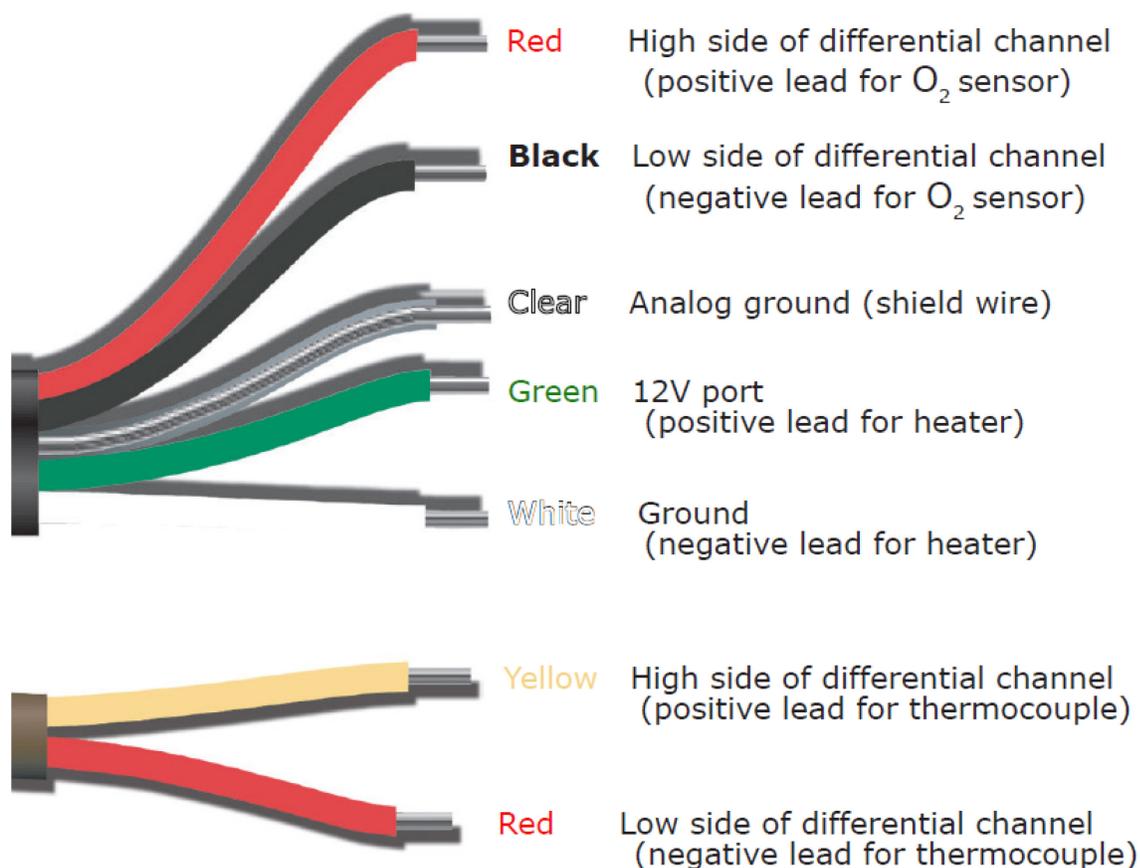
Connect the sensor to a measurement device (meter, datalogger, controller) capable of measuring and displaying or recording a millivolt (mV) signal (an input measurement range of approximately 0-60 mV or 0-15 mV is required to cover an oxygen range of 0-20.95 % for SO-100 series and SO-200 series sensors, respectively). In order to maximize measurement resolution and signal-to-noise ratio, the input range of the measurement device should closely match the output range of the oxygen sensor.

SO-110 and SO-210 (thermistor for sensor reference temperature measurement):



The oxygen sensor models with a thermistor for reference temperature measurement, SO-110 and SO-210, are recommended. Thermistors only use a single-ended datalogger channel (thermocouples use a differential channel) and do not require temperature measurement of the datalogger wiring panel (thermocouples require reference, or cold junction, temperature measurement, in addition to measurement of voltage from the thermocouple).

SO-120 and SO-220 Series (type-K thermocouple for sensor reference temperature measurement):



Absolute and Relative Gas Concentration:

Gas concentration is described in two ways, absolute and relative concentration. The ideal gas law yields absolute gas concentration, often expressed in quantity per volume [mol m⁻³] or partial pressure [kPa]:

$$PV = nRT \quad (1)$$

where P is pressure [Pa], V is volume [m³], n is gas quantity [mol], T is temperature [K], R is the ideal gas constant (8.314 J mol⁻¹ K⁻¹), and rearrangement of equation (1) to solve for n / V or P yields absolute gas concentration (in mol m⁻³ or kPa, respectively). However, a simple and common way to report concentration of a specific gas in a mixture is by expressing it relative to other gases in the mixture, as a fraction or percentage. For example, the amount of oxygen in the atmosphere, assuming a dry atmosphere (no water vapor), is 0.2095 kPa O₂ per kPa air, or 20.95 %. Atmospheric concentration of oxygen has remained constant for several hundred years at 20.95 %, and this percentage is the same at all elevations. However, absolute oxygen concentration does not remain constant (e.g., pressure decreases with elevation, thus absolute oxygen concentration decreases with elevation). Absolute oxygen concentration determines the rate of most biological and chemical processes, but relative oxygen concentration is often reported. This is analogous to measuring and reporting relative humidity when absolute humidity is what determines evaporation rates. Absolute and relative gas concentration measurements can be expressed using several different units.

Units Used to Describe Absolute and Relative Gas Concentration Measurements

Absolute Amount of Gas	Relative Amount of Gas
moles of O₂ per unit volume (e.g., moles per m ³ or moles per liter)	% O₂ in air (e.g., 20.95 % in ambient air)
mass of O₂ per unit volume (e.g., grams per liter; O ₂ has a mass of 32 g per mole)	mole fraction (e.g., moles of O ₂ per mole of air; 0.2095 mol O ₂ per mole of ambient air; this can also be expressed as 0.2095 kPa O ₂ per kPa air)
partial pressure (e.g., kilopascals [kPa])	

Sensor Calibration:

All Apogee oxygen sensors respond to absolute oxygen concentration in air, where common units of absolute gas concentration are partial pressure (e.g., kilopascals, kPa), mass per unit volume (e.g., grams per liter, g l⁻¹), and number of molecules per unit volume (e.g., moles per liter, mol l⁻¹). The absolute amount of oxygen in air is dependent on absolute (barometric) pressure and temperature, in addition to oxygen content of air. Therefore, Apogee oxygen sensors are not calibrated at the factory and must be calibrated by the user.

The output of Apogee oxygen sensors is a linear function of absolute oxygen concentration. A simple linear calibration is generally used to derive a calibration factor used to convert sensor output to relative oxygen concentration. The calibration factor (CF, in kPa O₂ mV⁻¹) is derived by dividing ambient oxygen partial pressure (21.23 kPa at sea level assuming standard pressure of 101.325 kPa) by the measured voltage output from the sensor under ambient conditions (in air or over water in a sealed chamber) minus the measured voltage output under conditions of zero oxygen (0 kPa O₂):

$$CF = \frac{0.2095P_B}{mV_C - mV_0} \quad (2)$$

where P_B is barometric pressure [kPa], 0.2095 multiplied by P_B equals partial pressure of oxygen under ambient conditions [kPa], mV_C is sensor voltage output [mV] during calibration, mV₀ is sensor voltage output [mV] under zero oxygen (0 kPa O₂), and CF is a linear multiplier that converts voltage measurements from the sensor to partial pressure of oxygen [kPa] using the equation:

$$O_2 = CF \cdot mV_M - \text{Offset} \quad (3)$$

where mV_M is measured voltage output [mV] and Offset is derived by multiplying CF by mV₀. The voltage output during calibration, mV_C, should be measured in a well-ventilated area. Do not breathe on the sensor, as exhaled breath has a much lower oxygen concentration than ambient air. If mV₀ is not measured, it can be estimated to be 3.0 mV for SO-100 series sensors and 0.30 mV for SO-200 series sensors. It is recommended that mV₀ be measured (in pure nitrogen gas) for applications where low values of oxygen (less than 10 kPa) will be measured. Precise measurements of hypoxic and anaerobic conditions can be made by making a periodic zero calibration of the sensor with ultra-pure nitrogen gas.

To convert sensor voltage output to partial pressure of oxygen (in kPa), multiply the measured voltage signal by the calibration factor, and then subtract the offset. For example, at sea level and 20.95 % O₂:

Calibration Factor [kPa O₂ per mV] * Sensor Output Signal [mV] - Offset [kPa] = Oxygen [kPa]

$$0.379 \quad * \quad 59.0 \quad - \quad 1.14 \quad = \quad 21.23$$

The calibration factor and offset are variable from sensor to sensor (those listed above are examples), and a sensor-specific calibration factor should be derived for each individual sensor. For routine oxygen measurements, the generic offset described above can be used. For measurements in air with less than 10 kPa (approximately 10 %) oxygen, a sensor-specific offset should be derived for each individual sensor.

Sensors can also be calibrated to measure relative oxygen concentration. The same procedure described for calibration to absolute oxygen is used, except ambient oxygen is set equal to 20.95 % (instead of 0.2095 multiplied by barometric pressure) to derive the calibration factor [% O₂ mV⁻¹]:

$$CF = \frac{20.95\%}{mV_C - mV_0} \quad (4)$$

where mV_C and mV_0 are as described above. The offset is also derived in the same manner, where mV_0 is multiplied by the calibration factor calculated from equation (4). Equation (3) is then used to produce relative oxygen measurements, when the calibration factor and offset derived from 20.95 % are used.

Changes in barometric pressure and temperature cause changes in absolute oxygen concentration, and as a result, changes in sensor signal output. This causes apparent changes in relative oxygen concentration, even though the relative amount of oxygen remains constant. Thus, barometric pressure and temperature corrections must be applied to relative oxygen measurements. Changes in absolute humidity (water vapor pressure of air) cause changes in absolute and relative oxygen concentration, as water vapor molecules displace and dilute oxygen molecules. Even though changes in water vapor content cause actual (not apparent) changes in relative oxygen concentration, water vapor effects are often corrected for to yield relative oxygen concentrations for a dry atmosphere.

Effect of Barometric Pressure on Oxygen Concentration:

The ideal gas law, equation (1), shows that absolute gas concentration increases by 0.987 % at sea level for every 1 kPa increase in pressure (1 kPa / 101.325 kPa = 0.00987). For a sensor that measures absolute gas concentration, but is calibrated to read out in relative units, a 1 kPa pressure increase at sea level results in an apparent oxygen increase of 0.207 % (0.00987 * 20.95 % = 0.207 %) and an apparent relative oxygen concentration of 21.157 %. Relative gas concentration didn't really increase, but absolute concentration, which is what sensors measure, did change. This shows up as an apparent change in relative concentration.

Due to lower barometric pressure at higher elevations, the percentage increase in absolute gas concentration per kPa increases with elevation. For example, at an elevation of 1378 m (Logan, Utah), barometric pressure is approximately 86 kPa and absolute gas concentration increases by 1.16 % for every 1 kPa increase in pressure (1 kPa / 86 kPa = 0.0116). Again, for a sensor that measures absolute gas concentration, but is calibrated to read out in relative units, this results in an apparent oxygen increase. In this example, 0.247 % for every 1 kPa increase in barometric pressure (0.0118 * 20.95 % = 0.243 %) and an apparent relative oxygen concentration of 21.193 %.

A barometric pressure correction should be applied to all oxygen sensors that are calibrated to read relative oxygen concentration. The equation to correct relative oxygen measurements for barometric pressure at any elevation is:

$$O_2 = O_{2M} \left(\frac{P_C}{P_M} \right) \quad (5)$$

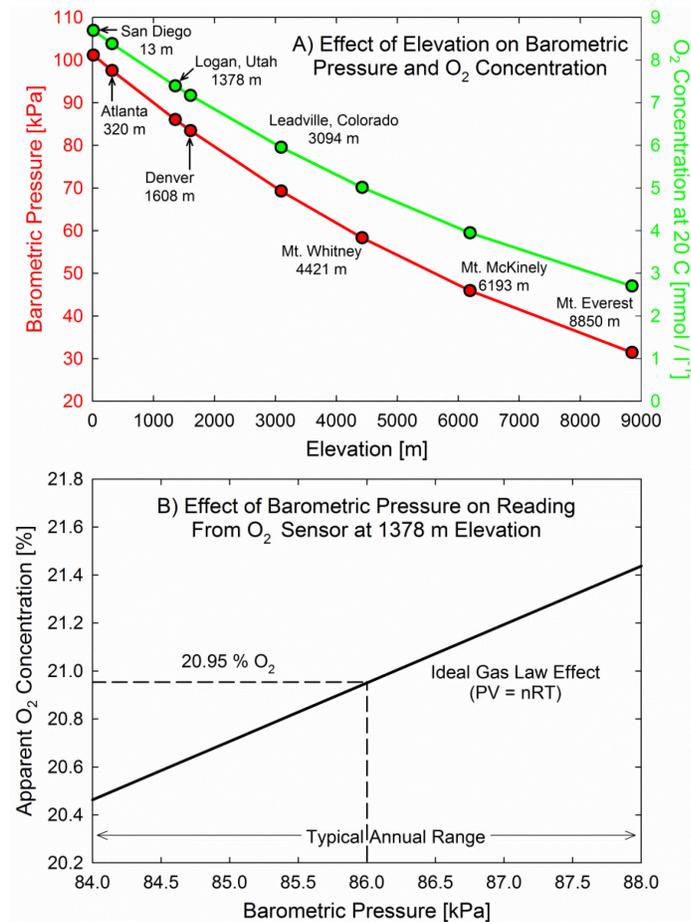
where O_{2M} is measured oxygen concentration [%] (apparent oxygen concentration), P_C is barometric pressure [kPa] at the time of calibration, and P_M is barometric pressure [kPa] at the time of the current measurement. Approximate barometric pressure (P_B , in kPa) for a given elevation is calculated from:

$$P_B = 101.325 - 101.325 \left[1 - \left(1 - \frac{E}{44307.69231} \right)^{5.25328} \right] \quad (6)$$

where E is elevation [m]. In order to make a barometric pressure correction on gas measurements, it must be continuously measured as it changes over time (see Apogee webpage for a barometric pressure sensor that can be used for continuous measurements of barometric pressure:

<http://www.apogeeinstruments.com/barometric-pressure/>). The typical annual barometric pressure range is approximately 4 kPa, or the average pressure for a given elevation +/- 2 kPa.

The apparent effect of barometric pressure on relative oxygen measurements, based on calculations from equation (5), is plotted in the figure below for 1378 m elevation to show the significance of measuring and correcting for barometric pressure. If not accounted for, barometric pressure fluctuations show up in oxygen measurements as a change in relative oxygen concentration because sensors respond to absolute oxygen concentration, but are generally calibrated to read out in relative units.



A) Barometric pressure and absolute oxygen concentration at 20 C as a function of elevation. Equation (6) was used to calculate barometric pressure. B) Effect of barometric pressure on apparent relative oxygen concentration. Oxygen sensors respond to absolute oxygen concentration, but are often calibrated to yield relative oxygen concentration. As barometric pressure fluctuates, absolute oxygen concentration, and thus oxygen sensor output, fluctuates with it, producing an apparent change in relative oxygen concentration if this pressure effect is not accounted for. It is assumed the sensor was calibrated at 86 kPa, and the solid line shows how the apparent relative oxygen concentration is dependent on barometric pressure.

Effect of Temperature on Oxygen Concentration:

The ideal gas law, equation (1), shows that absolute gas concentration decreases by 0.341 % for a 1 C increase in temperature from 20 C (1 K / 293 K = 0.00341). For a sensor that measures absolute gas concentration, but is calibrated to read out in relative units, a 1 C temperature increase from 20 C results in an apparent decrease of 0.0714 % O₂ (0.341 % * 0.2095 = 0.0714 %) and a relative oxygen concentration of 20.878 %. As with barometric pressure, to obtain accurate oxygen measurements with a sensor that is calibrated to read relative oxygen concentration, a correction should be applied to compensate for temperature effects. The equation to correct relative oxygen measurements in air for temperature effects is:

$$O_2 = O_{2M} \left(\frac{T_M}{T_C} \right) \quad (7)$$

where O_{2M} is as given above, T_C is air temperature [K] at calibration, and T_M is air temperature [K] at the time of measurement (note that temperatures in equation (7) must be in K). The effects of temperature on relative oxygen concentration measurements, based on calculations from equation (7), are plotted in the figure below to show the significance of measuring and correcting for temperature. If not accounted for, temperature fluctuations show up in the measurement as an apparent change in relative oxygen concentration because sensors respond to absolute oxygen concentration, but are calibrated to read out in relative units.

Sensor Response to Temperature

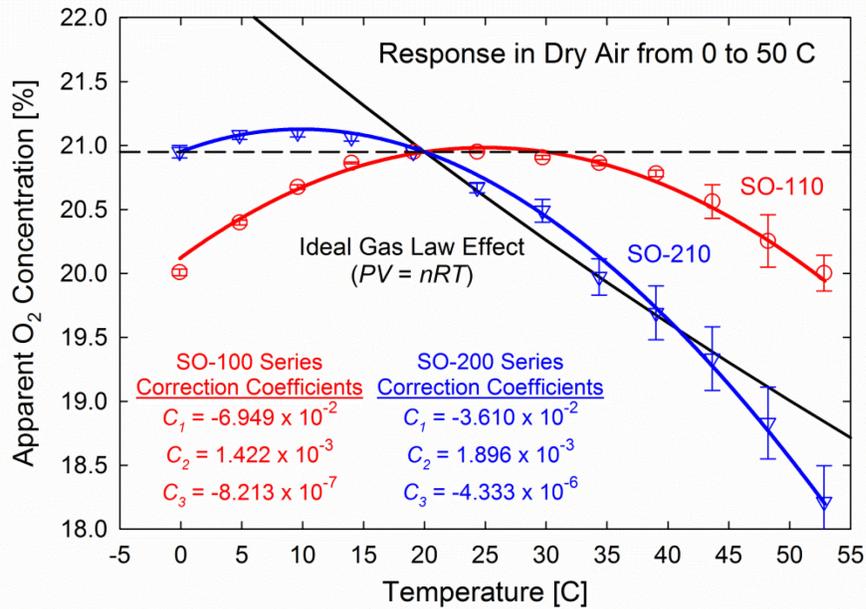
In practice, equation (7) does not accurately correct for temperature effects because in addition to the ideal gas law temperature effect, sensor electronics are affected by temperature. The combination of these two effects on Apogee oxygen sensors (SO-100 and SO-200 series) was determined from measurements in dry air across a wide temperature range by plotting pressure-corrected apparent oxygen concentration (i.e., measured oxygen concentration before temperature correction was applied) versus measured sensor temperature (T_S). Neither of the sensor models (SO-100 or SO-200 series) follows the ideal gas law response, and thus an empirical correction derived from measured data must be applied to account for both the ideal gas law and sensor electronics responses:

$$O_2 = O_{2M} + C_3 T_S^3 + C_2 T_S^2 + C_1 T_S + C_0 \quad (8)$$

where T_S is measured sensor temperature [C] (Apogee oxygen sensors come with a thermistor or type-K thermocouple temperature reference sensor); coefficients C₃, C₂, and C₁ are listed in the figure below for both SO-100 and SO-200 series sensors; and C₀ is the offset coefficient calculated from measured temperature at calibration (T_C, in C):

$$C_0 = -(C_3 T_C^3 + C_2 T_C^2 + C_1 T_C). \quad (9)$$

The temperature effect on sensor electronics is slightly variable from sensor to sensor, thus coefficients derived (average of three replicate sensors; error bars representing two standard deviations are shown in figure) may not yield the most accurate temperature correction for all sensors of the same model.



Empirically-measured temperature responses of SO-100 and SO-200 series oxygen sensors, with third order polynomials fit to data, compared to the theoretical temperature response calculated from the ideal gas law, equation (1). The difference between theoretical and measured responses is due to a temperature effect on sensor electronics. The polynomial coefficients used to correct for the temperature response with equation (8) are listed. An offset coefficient (C_0) is not listed because it is dependent on temperature at calibration. It is calculated with equation (9). Sensors were calibrated at 20 C. As with barometric pressure, absolute oxygen concentration, and thus oxygen sensor output, varies with temperature. As temperature changes, relative oxygen concentration remains constant at 20.95 %, but an apparent oxygen change is measured if the temperature correction is not applied to relative measurements.

Effect of Humidity on Oxygen Concentration:

As absolute humidity in the atmosphere increases, water vapor molecules displace and dilute other gas molecules. This causes the signal output of a gas sensor to decrease. The water vapor effect on relative oxygen concentration as a function of relative humidity (RH) and at a constant temperature is a linear decrease with increasing RH, as shown in the figure below. Conversely, the effect as a function of temperature at constant RH is a curvilinear decrease with increasing temperature, essentially the inverse of the slope of vapor pressure curves from a psychrometric chart. Even though water vapor molecules dilute and displace oxygen molecules, and thus cause an actual and not an apparent decrease in relative oxygen concentration, humidity effects are often accounted for to yield relative oxygen concentrations for a dry atmosphere. The equation to correct for humidity effects is:

$$O_2 = O_{2M} \left(\frac{P_C + (e_{AM} - e_{AC})}{P_C} \right) \quad (10)$$

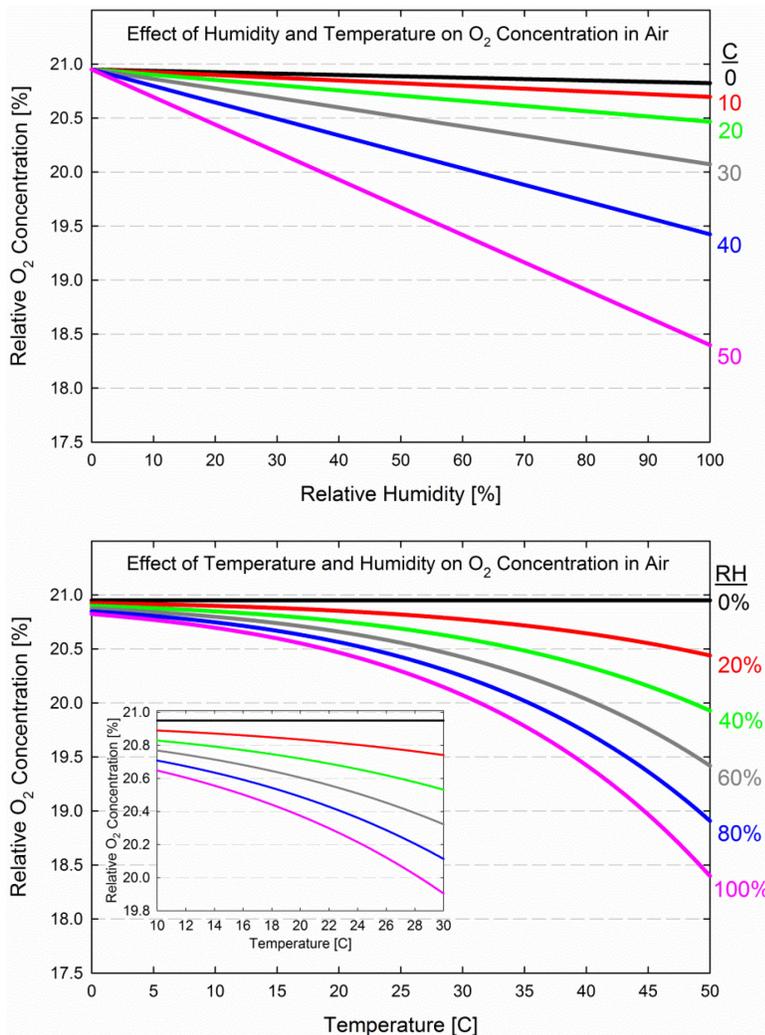
where P_C is barometric pressure at calibration [kPa], e_{AM} is vapor pressure [kPa] of air at the time of measurement, and e_{AC} is vapor pressure [kPa] of air at calibration. Vapor pressures in equation (10) are calculated from:

$$e_A = e_S \left(\frac{RH}{100} \right) \quad (11)$$

where RH is in % and e_S is saturation vapor pressure [kPa] of air calculated from air temperature (T_A , in C):

$$e_S = 0.61121 \exp \left(\frac{T_A \left(\frac{18.678 - T_A}{234.5} \right)}{257.14 + T_A} \right) \quad (12)$$

In soil environments relative humidity is generally between 99 and 100 %, unless the soil is extremely dry (below the permanent wilting point of -1,500 kPa). Thus, the water vapor effect can be accounted for as a function of temperature by correcting oxygen measurements based on the shape of the curve for 100 % RH in the graph below.



A) Relative humidity effects on relative oxygen concentration shown as a function of relative humidity at temperatures increments of 10 C and B) as a function of temperature at relative humidity increments of 20 %. The air in soil is typically always saturated with water vapor (100 % relative humidity) unless the soil is very dry.

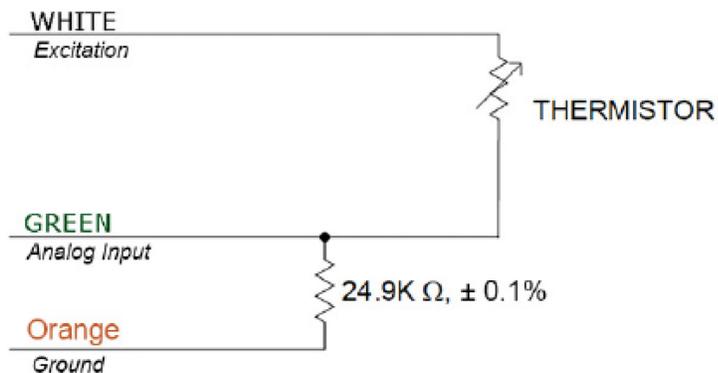
As with temperature, humidity also causes a slight effect on the sensor electronics. For measurements in soil or saturated air (100 % relative humidity), it is recommended that Apogee oxygen sensors are calibrated in conditions where relative humidity is 100 %. A simple way to accomplish this is to mount the sensor in a sealed chamber over water, with ambient air filling the headspace, as shown below.



Apogee oxygen sensor mounted in a sealed chamber over water. For measurements in environments where relative humidity is 100 %, sensors should be calibrated in conditions where relative humidity is 100 % in order to account for any humidity effects on sensor electronics.

Temperature Measurement with Internal Thermistor:

Measurement devices (e.g., datalogger, controller) do not measure resistance directly, but determine resistance from a half-bridge measurement, where an excitation voltage is input across the thermistor and an output voltage is measured across the bridge resistor.



An excitation voltage of 2.5 V DC is recommended to minimize self-heating and current drain, while still maintaining adequate measurement sensitivity (mV output from thermistor per C). However, other

excitation voltages can be used. Decreasing the excitation voltage will decrease self-heating and current drain, but will also decrease thermistor measurement sensitivity. Increasing the excitation voltage will increase thermistor measurement sensitivity, but will also increase self-heating and current drain.

The internal thermistor provides a temperature reference for correction of relative oxygen measurements for temperature changes. Resistance of the thermistor changes with temperature. Thermistor resistance (R_T , in Ω) is measured with a half-bridge measurement, requiring an excitation voltage input (V_{EX}) and a measurement of output voltage (V_{OUT}):

$$R_T = 24900 \left(\frac{V_{EX}}{V_{OUT}} - 1 \right) \quad (13)$$

where 24900 is the resistance of the bridge resistor in Ω . From resistance, temperature (T_K , in Kelvin) is calculated with the Steinhart-Hart equation and thermistor specific coefficients:

$$T_K = \frac{1}{A + B \ln(R_T) + C (\ln(R_T))^3} \quad (14)$$

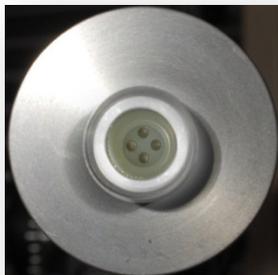
where $A = 1.129241 \times 10^{-3}$, $B = 2.341077 \times 10^{-4}$, and $C = 8.775468 \times 10^{-8}$.

Heating Sensor with Internal Heater:

All Apogee oxygen sensors are equipped with an internal resistance heater. The heater is designed to maintain the temperature of the sensing element at approximately 2 C above ambient temperature in condensing (100 % relative humidity) environments (e.g., soil). Heating the sensing element keeps condensation from forming on the membrane, which would block the oxygen diffusion path and result in erroneous measurements. To operate the heater, apply continuous 12 V DC across the yellow (positive) and blue (negative) wires.

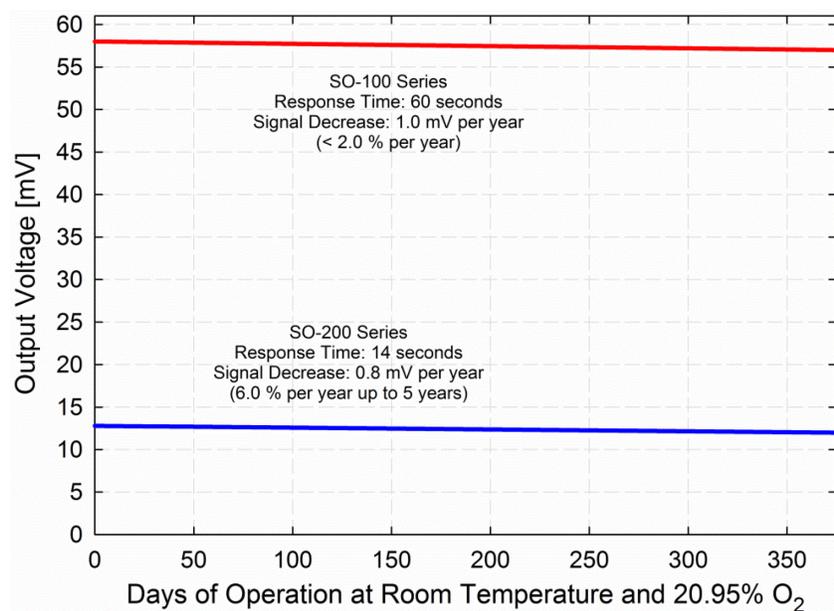
MAINTENANCE AND RECALIBRATION

Visual inspection of the Teflon membrane should be made periodically to verify that the oxygen path is free from obstruction, as shown below. Avoid placing sharp objects inside the sensor opening, as the membrane can easily be punctured.



Life expectancy of SO-100 and SO-200 series sensors is approximately ten and five years of continuous use in 20.95 % oxygen, respectively. Lifetime can be lengthened by storing sensors in cold temperatures (e.g., fridge or freezer) when not in use.

Sensor recalibration can be conducted periodically and should be determined by the level of measurement accuracy required for the application. Sensor signal decrease over one year when exposed to 20.95 % oxygen is shown in the figure below. SO-100 and SO-200 series sensors decrease by approximately 1 and 0.8 mV per year, respectively, or approximately 2 and 6 % of signal output at 20.95 % oxygen. This signal decrease yields increases in calibration factor of approximately 2 and 6 % for SO-100 and SO-200 series sensors, respectively.



Long-term stability (output voltage decrease over time) of Apogee SO-100 and SO-200 series oxygen sensors. The response time and signal decrease for both series are also listed.

TROUBLESHOOTING AND CUSTOMER SUPPORT

Independent Verification of Functionality:

The oxygen sensing element inside Apogee SO-100 and SO-200 series oxygen sensors is a self-powered device that outputs a voltage signal proportional to partial pressure of gaseous oxygen. A quick and easy check of sensor functionality can be determined using a voltmeter with millivolt (mV) resolution. Connect the positive lead of the voltmeter to the red wire from the sensor and the negative lead (or common) to the black wire from the sensor. SO-100 series sensors should read approximately 60 mV at sea level in ambient air (20.95 % O₂), and SO-200 series sensors should read approximately 12 mV at sea level in ambient air. These voltages will decrease by approximately 1 % per 100 meters of elevation increase above sea level.

The thermistor inside Apogee SO-110 and SO-210 oxygen sensors yields a resistance proportional to temperature. A quick and easy check of thermistor functionality can be accomplished with an ohmmeter. Connect the lead wires of the ohmmeter to the green and white wires from the sensor. The resistance should read 10 k Ω at 25 C. If the sensor temperature is less than 25 C, the resistance will be higher. If the sensor temperature is greater than 25 C, the resistance will be lower. Connect the lead wires of the ohmmeter to the green and orange wires from the sensor. The resistance should read 24.9 k Ω , and should not vary. Connect the lead wires of the ohmmeter to the white and orange wires from the sensor. The resistance should be the sum of the resistances measured across the green and white wires, and green and orange wires (e.g., 10 k Ω plus 24.9 k Ω at 25 C).

The thermocouple inside Apogee SO-120 and SO-220 oxygen sensors yields a voltage proportional to the temperature difference between the hot (thermocouple tip inside sensor) and cold (thermocouple pigtail lead wires) junctions. A quick and easy check of thermocouple functionality can be determined using a voltmeter in continuity mode. Connect the leads of the voltmeter to the leads of the thermocouple (polarity does not matter), and test for continuity. Continuity indicates a functional thermocouple.

Compatible Measurement Devices (Dataloggers/Controllers/Meters):

SO-100 series oxygen sensors have a calibration factor of approximately 0.39 kPa O₂ per mV, yielding a sensitivity of approximately 2.6 mV per kPa O₂. SO-200 series oxygen sensors have a calibration factor of approximately 1.7 kPa O₂ per mV, yielding a sensitivity of approximately 0.6 mV per kPa O₂. Thus, a compatible measurement device (e.g., datalogger or controller) should have resolution of at least 0.02 and 0.006 mV, in order to provide pressure resolution of 0.01 kPa O₂ for the SO-100 and SO-200 series oxygen sensors, respectively.

The signal output range of SO-100 and SO-200 series oxygen sensors is approximately 3 to 250 mV and 0.3 to 60 mV, respectively. A compatible measurement device should also have a full scale range spanning the sensor signal output range.

An example datalogger program for Campbell Scientific dataloggers can be found on the Apogee webpage at <http://www.apogeeinstruments.com/content/Oxygen-Sensor.CR1>.

Modifying Cable Length:

When the sensor is connected to a measurement device with high input impedance, sensor output signals are not changed by splicing on additional cable in the field. Tests have shown that if the input impedance of the measurement device is 1 mega-ohm or higher, then there is negligible effect on SO-100 and SO-200 series oxygen sensor calibrations, even after adding up to 100 m of cable. See Apogee webpage for details on how to extend sensor cable length (<http://www.apogeeinstruments.com/how-to-make-a-weatherproof-cable-splice/>). For cable extensions, shielded, twisted pair cable is recommended, in order to minimize electromagnetic interference. This is particularly important for long lead lengths in electromagnetically noisy environments.

RETURN AND WARRANTY POLICY

RETURN POLICY

Apogee Instruments will accept returns within 30 days of purchase as long as the product is in new condition (to be determined by Apogee). Returns are subject to a 10 % restocking fee.

WARRANTY POLICY

What is Covered

All products manufactured by Apogee Instruments are warranted to be free from defects in materials and craftsmanship for a period of four (4) years from the date of shipment from our factory. To be considered for warranty coverage an item must be evaluated either at our factory or by an authorized distributor.

Products not manufactured by Apogee (spectroradiometers, chlorophyll content meters) are covered for a period of one (1) year.

What is Not Covered

The customer is responsible for all costs associated with the removal, reinstallation, and shipping of suspected warranty items to our factory.

The warranty does not cover equipment that has been damaged due to the following conditions:

1. Improper installation or abuse.
2. Operation of the instrument outside of its specified operating range.
3. Natural occurrences such as lightning, fire, etc.
4. Unauthorized modification.
5. Improper or unauthorized repair.

Please note that nominal accuracy drift is normal over time. Routine recalibration of sensors/meters is considered part of proper maintenance and is not covered under warranty.

Who is Covered

This warranty covers the original purchaser of the product or other party who may own it during the warranty period.

What We Will Do

At no charge we will:

1. Either repair or replace (at our discretion) the item under warranty.
2. Ship the item back to the customer by the carrier of our choice.

Different or expedited shipping methods will be at the customer's expense.

How To Return An Item

1. Please do not send any products back to Apogee Instruments until you have received a Return Merchandise Authorization (RMA) number from our technical support department by calling (435) 792-4700 or by submitting an online RMA form at www.apogeeinstruments.com/tech-support-recalibration-repairs/. We will use your RMA number for tracking of the service item.
2. Send all RMA sensors and meters back in the following condition: Clean the sensor's exterior and cord. Do not modify the sensors or wires, including splicing, cutting wire leads, etc. If a connector has been attached to the cable end, please include the mating connector – otherwise the sensor connector will be removed in order to complete the repair/recalibration.

3. Please write the RMA number on the outside of the shipping container.
4. Return the item with freight pre-paid and fully insured to our factory address shown below. We are not responsible for any costs associated with the transportation of products across international borders.
5. Upon receipt, Apogee Instruments will determine the cause of failure. If the product is found to be defective in terms of operation to the published specifications due to a failure of product materials or craftsmanship, Apogee Instruments will repair or replace the items free of charge. If it is determined that your product is not covered under warranty, you will be informed and given an estimated repair/replacement cost.

Apogee Instruments, Inc.
721 West 1800 North Logan, UT
84321, USA

OTHER TERMS

The available remedy of defects under this warranty is for the repair or replacement of the original product, and Apogee Instruments is not responsible for any direct, indirect, incidental, or consequential damages, including but not limited to loss of income, loss of revenue, loss of profit, loss of wages, loss of time, loss of sales, accrual of debts or expenses, injury to personal property, or injury to any person or any other type of damage or loss.

This limited warranty and any disputes arising out of or in connection with this limited warranty ("Disputes") shall be governed by the laws of the State of Utah, USA, excluding conflicts of law principles and excluding the Convention for the International Sale of Goods. The courts located in the State of Utah, USA, shall have exclusive jurisdiction over any Disputes.

This limited warranty gives you specific legal rights, and you may also have other rights, which vary from state to state and jurisdiction to jurisdiction, and which shall not be affected by this limited warranty. This warranty extends only to you and cannot be transferred or assigned. If any provision of this limited warranty is unlawful, void or unenforceable, that provision shall be deemed severable and shall not affect any remaining provisions. In case of any inconsistency between the English and other versions of this limited warranty, the English version shall prevail.

This warranty cannot be changed, assumed, or amended by any other person or agreement.

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